

the scientist is faced with two additional problems: opposition by those who believe animals should not be used in experiments; and the inherent one posed by the introduction of a living organism, a variable with many unknown characteristics, into an experiment.

**Lab Animals Are Protected**

Stringent safeguards are in effect that protect the laboratory animal. Anesthesia must be administered in experiments causing pain. In certain circumstances where recovery of the animal is not possible, it must be destroyed painlessly and quickly. Other rulings exist to ensure adequate food, etc.

Regulations also exist concerning the procurement of animals for research purposes. The breeding of laboratory animals specifically for research has benefited the scientist in two ways: he receives an animal from a legitimate source; the animal comes with a more or less well-defined genetic makeup, and is thus less of an "unknown" in the experiment.

Groups such as the Animal Care Panel and the National Society for Medical Research are continually working to improve the condition of laboratory animals and to make known the problems faced by the biologist in working with a living creature.

Experimentation is indispensable to the biologist just as it is to the rocket engineer or the physicist. Such research can benefit all life, whether it is the mouse or man.

Science News Letter, February 7, 1959

**METEOROLOGY**

**Suggest "Blue" Salt Aids in Making Rain**

► "BLUE" SALT in the atmosphere may play an important role in making rain, a Canadian scientist told the American Meteorological Society meeting in New York.

Dr. Henry M. Papee of the division of applied chemistry, National Research Council, Ottawa, said laboratory studies showed "blue" salt adsorbs water vapor much more readily than normal salt. The "blue" or activated salt, he suggested, is formed in the atmosphere by "radio-chemical excitation."

He prepared the "blue" salt by treating microcrystals of sodium chloride with radio frequency waves. The surface of the treated salt decreases much more slowly than that of normal sodium chloride when exposed to water vapor, Dr. Papee reported.

The blue color is due to colloidal sodium on the outer surface, making the activated salt "remarkably stable" at room temperatures.

Whether rain is formed by the melting of ice crystals or the growth of tiny drops of water vapor, there must be a minute particle around which the crystal or the raindrop forms. Weathermen call such particles condensation nuclei.

Dr. Papee reported that the treated salt would provide a "much more efficient condensation nucleus" than ordinary salt.

Science News Letter, February 7, 1959

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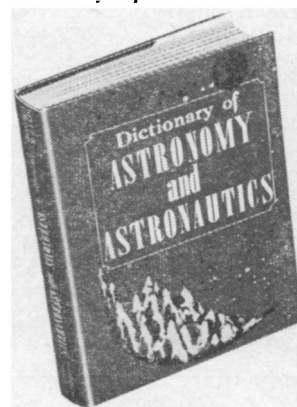
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